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Section 2-1: The Nature of Matter The subatomic particles that make up atoms are protons, neutrons, and electrons. Because they have the same number of protons, all isotopes of an element have the same chemical properties. The main types of chemical bonds are covalent bonds and ionic bonds.

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the building blocks of matter. proton.
positively charged part of an atom in the
nucleus. neutron. part of atom with no
charge in the nucleus. electron.
negatively charged particle of an atom
that orbits around the nucleus and is
1/1840 the size of a proton.

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Chapter 2: The Chemistry of Life
The Nature of Matter
Atoms are the smallest, indivisible particles of matter
First theorized by Democritus about 2500 years ago (ya)
Atoms made of two parts that consist of a combination of three different subatomic particles
Nucleus is the "core" of an atom
Protons have a positive charge
Neutrons have a neutral charge
Electron cloud
Divided into orbitals (layers)
First layer is smallest and contains maximum of 2 electrons
Second layer contains maximum of 8 ...

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text pp 190-192 Microscopes. Chapter 2.1-2.2 The Chemistry of Life. Chapter 3 The Biosphere

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