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Chapter 6 Solutions Engineering and Chemical Thermodynamics Wyatt Tenhaeff Milo Koretsky Department of Chemical Engineering Oregon State University 2 6.1 (a) The Clausius-Clapeyron equation: $dP_i^{\text{sat}}/dT = h_i^{\text{vap}}/RT^2$ or $\ln P_i^{\text{sat}}(101 \text{ kPa}) = \frac{h_i^{\text{vap}}}{R} \left(\frac{1}{T} - \frac{1}{1373 \text{ K}} \right) + \ln P_i^{\text{sat}}(101 \text{ kPa})$ so $P_i^{\text{sat}} = 101 \text{ kPa} \exp \left(\frac{h_i^{\text{vap}}}{R} \left(\frac{1}{T} - \frac{1}{1373 \text{ K}} \right) \right)$

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Chapter 6 2 If the fluid is in thermodynamic equilibrium any thermodynamic variable for a pure substance, like pure water, can be written in terms of any two other thermodynamic variables , i.e. $p=p(\rho,T)$ (6.1.1) where the functional relationship in depends on the substance.

Chapter 6 Thermodynamics and the Equations of Motion

Chapter 6: Solution Thermodynamics and Principles of Phase Equilibria In all the preceding chapters we have focused primarily on thermodynamic systems comprising pure substances. However, in all of nature, mixtures are ubiquitous.

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