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A.P. Chemistry Practice
Test - Ch. 13:

Equilibrium Name _____
MULTIPLE CHOICE.

Choose the one
alternative that best
completes the
statement or answers
the question. 1) At
equilibrium, _____.
A) the rates of the
forward and reverse
reactions are equal
B) the rate constants of
the forward and

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reverse reactions are
equal

**A.P. Chemistry
Practice Test - Ch.
13: Equilibrium ...**

AP Chemistry Chapter
13. component.
solvent. solutes.
aqueous solutions.
each of the substances
in a solution. normally
the component present
in the greatest amount.
components within
solution that are in
lesser amounts than

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oth.... solutions that
contain water as a
solvent and a gas,
liquid, o....

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AP Chemistry

Retest/Brownie Points
(Chapter 13) Multiple
Choice (20%) 1) Please
write the equilibrium
expression for this
reaction. 3 A (aq) + 3
D (s) \leftrightarrow 4 C (l) + B (g)

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+ 2 E (s) 2) Please consider the gas phase reaction: $2 \text{NO} + \text{Br}_2 \leftrightarrow 2 \text{BrNO}$ $\Delta H = -345 \text{ kJ/mol}$. Which one would increase the yield of NO?

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and ...**

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(Chapter 12) Multiple
Choice (40%) AP
Chemistry Test
(Chapter 12) Multiple
Choice (40%) 1) Which
of the following is a &
Chapter 13 Chapter 11:
*Below is a phase
diagram for carbon

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dioxide Use it to
answer questions 1
and 2* 1At what

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Answers . AP Question .
Ch 14 and 19: Ch 14
Ch19 Practice Keys .
Ch 15 pH Calculation
Answers. Chapter 15

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Acid Base Practice Test
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15 Handout Keys . Ch
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Baker, Mrs.
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Chapter 13 - Chemical Equilibrium . Intro . A. Chemical Equilibrium 1. The state where the concentrations of all reactants and products remain constant with time 2. All reactions carried out in a closed vessel will reach equilibrium a. If little product is formed, equilibrium lies far to the left b.

**Chapter 13 -
Chemical**

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**Equilibrium -
ScienceGeek.net**

AP Chemistry: Chapter
13--Equilibrium

Chapter 13 Test

Review Know how to
write K_c , K_p

expressions (*note that
solids and liquids not
included; K_c in [], and

K_p in atm) Know how
to calculate K_c , K_p

when: Initial

concentrations (or
pressures) are known

Changes are known

Equilibrium

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concentrations (or pressures) are known
*Note for each of the above, ICE diagrams are very helpful!!!

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Chapter
13--Equilibrium
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investigations as they explore the four Big Ideas: scale, proportion, and quantity; structure and properties of substances; transformations; and energy.

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Practice Quiz

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Chapter 13 - Chemical Equilibrium | CourseNotes

This quiz is to review for the Chapter 13 test.

Chem II Chapter 13 -

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ProProfs Quiz

Chapter 13. AP
Chemistry. Notes:
Chemical Equilibrium.
Many reactions do not
run to completion. But,
reaction concentrations
may cease to change.
The system contains a
mixture of reactants
and products. A.
chemical equilibrium.
has been reached.

**AP Chemistry Notes:
Chapter 15 Chemical
Equilibrium**

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1 AP Chemistry Name:

_____ 11/07/17 Chapter

13 MC 1. $2 XY(g) \leftrightarrow X_2$

$(g) + Y_2(g)$ $K_p = 230$

A certain gas, $XY(g)$,

decomposes as

represented by the

equation above. A

sample of each of the

three gases is put in a

previously evacuated

container. The initial

partial pressures of the

gases are shown in the

table below.

Chapter 13 MC 2017

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**Chapter 14 Chemical
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1 Test Review. Days
1-10. Assignments.
Stoichiometry Review.
Test Review. ... Key -
Chapter 13 and 14 Test
Review. KEY - Chapter
13/14 Multiple Choice
Rev. Due Date Lab Due
12/12/18. Chapter 15 -
Equilibrium .
Assignments .

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